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Chemistry Vocab

Chapter 10. Avogadro's

Number. mole. molar

mass. Empirical

Formula. 6.02×10^{23} ;

counts extremely small

particles. an SI base

unit used to measure

the quantity of matter.

Mass in grams of one

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mole of any pure substance. Formula with smallest whole-number ratio of the elements.

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ANSWER KEY 1 REVIEW

QUESTIONS Chapter 10

1. Draw Lewis structures and determine the molecular geometry of each molecule or ion

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shown below: A) ClO_2 2
A 20 electrons Bent
(from tetrahedral) B)
 ICl_3 3 28 electrons T-
shaped (from trigonal
bipyramidal) C) TeF_4 4
34 electrons See-saw
(from trigonal
bipyramidal) 2.

Chemistry 101

ANSWER KEY -

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CHAPTER 10 REVIEW

States of Matter

SECTION 4 SHORT

ANSWER Answer the

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following questions in the space provided. 1. a When a substance in a closed system undergoes a phase change and the system reaches equilibrium, (a) the two opposing changes occur at equal rates. (b) there are no more phase changes. (c) one phase change predominates.

**10 States of Matter -
Ms. Agostine's
Chemistry Page**

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10.2 Mole to Mole & Mole to Volume

Relationships *For all the problems on this page, first find the molar mass of the

compound. 1. What is the mass of 9.45 mol of aluminum oxide?

$2\text{Al}_2\text{O}_3$

63,55 g/mol - 2. What is the mass of 4.52×10^{-3} mol of ethylbenzene, $\text{C}_6\text{H}_5\text{CH}_2\text{CH}_3$?

*Ethylbenzene is a hydrocarbon that is produced by burning

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Prentice Hall Chemistry
Chapter 10. Avogadro's
number. Empirical
Formula. Molar Mass.
Molar Volume. number
of particles in one mole
of a pure substance
(element o.... Formula
that shows the lowest
whole-number ratio of
the atoms.... The mass
of one mole of an
element. Found on the

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periodic tabl....

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Chemistry Chapter 10

Test Review Multiple

Choice Identify the

choice that best

completes the

statement or answers

the question. ____ 1.

According to the

kinetic-molecular

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Chapter 10 Review

theory, particles of matter a. are in constant motion. c. have different colors. b. have different shapes. d. are always fluid. ____ 2.

Chemistry Chapter 10 Test Review - Mr. Hoge's Science

Chapter 10 - Organic Chemistry Review

DRAFT. 11th - 12th grade. 91 times.

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months ago. kallen. 0.

Save. Edit. Edit.

Chapter 10 - Organic
Chemistry Review

DRAFT. ... answer

choices . CH 3 CH 2

CHO. CH 3 COCH 3. CH

3 CH 2 COOH. CH 3

COOCH 3. Tags:

Question 2 . SURVEY .

30 seconds . Q.

Chapter 10 - Organic

Chemistry Review

Quiz - Quizizz

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Chapter 10 Review

chapter assumes that you can already perform the tasks listed below. You can test your readiness to proceed by answering the Review Questions at the end of the chapter. This might also be a good time to read the Chapter Objectives, which precede the Review Questions.

1.09×10^4
kg P g P mol P mol P
40 10 ...

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Chapter 10 Review

**Chapter 10 Chemical
Calculations and
equations**

Name Date Class.

CHAPTER 10 REVIEW.

States of Matter.

SECTION 1. SHORT

ANSWERS Answer the following questions in the space provided. 1.

Identify whether the descriptions below describe an ideal gas or a real gas. a. The gas will not condense because the molecules do not attract each

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other.

**10 States of Matter -
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PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 4. a. Use five to six data points from Appendix Table A-8 in the text to sketch the curve for water vapor's partial pressure versus

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temperature on the
graph provided below.

30 25 20 15 10 No 10

torr MODERN

CHEMISTRY b. c. 10 15

20 25 30 Temperature

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