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undergraduate work at Iowa State University, completing her bachelor's degree and Master's degree in inorganic chemistry at the University of South Florida in Tampa. She earned her Ph.D. in analytical chemistry at the University of Idaho.

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Which should have the greater second ionization energy ( $IE_2$ )? Worked Example 4.4 Solution  $IE_1(\text{Mg}) > IE_1(\text{Na})$  because Mg is to the right of Na in the periodic table (i.e., Mg has the greater  $Z_{\text{eff}}$ , so it is more difficult to remove its electron).  $IE_2(\text{Na}) > IE_2(\text{Mg})$  because the second ionization of Mg removes a valence

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