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This hydrocarbon is called "2,3-pentadiene" because there are 2 (di) double bonds (ene) after 2 and 3 (2,3) of a 5 carbon line (penta). It is named the same way for triple bonds, but with an yne instead of ene. Next are isomers. Isomers are compounds

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with the same molecular formula but different molecular structures.

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\ Chapter 22 Hydrocarbon Compounds. Chapter 22 Hydrocarbon Compounds. Flashcard maker : Paula Corcoran. How is the number of valence electrons in carbon atoms related to the bonds that carbon atoms form? because carbon has four valence electrons, a carbon atom always forms four covalent bonds.

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Chapter 22 "Hydrocarbon Compounds". Use these activities to learn the vocabulary and major concepts presented in the chapter. molecules in which the atoms are joined in the same order, but the arrangements of the atoms in space are different.

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22.1 Hydrocarbons. • Because carbon has four valence electrons, carbon atoms always form four covalent bonds. • The carbon atoms in an alkane can be arranged in a straight chain or in a chain that has branches. • Molecules of hydrocarbons, such as alkanes, are nonpolar molecules. 22.2 Unsaturated Hydrocarbons.

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