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2 . 12. Chemistry 11

Calculations Practice

Test # 1 . 13. Instructional Fair

Chemistry 11

Calculations Practice

Test # 2 . Molarity

Worksheet # 1 . 1.

15.8 g of KCl is

dissolved in 225 mL of

water. Calculate the

molarity.

Molarity Worksheet

1

Dilutions Worksheet 1)

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If I add 25 mL of water to 125 mL of a 0.15 M NaOH solution, what will the molarity of the diluted solution be? 2)

If I add water to 100 mL of a 0.15 M NaOH solution until the final volume is 150 mL, what will the molarity of the diluted solution be? 3) How much 0.05 M HCl solution can be made by diluting 250 mL of 10 M HCl?

Dilutions Worksheet

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Dilution Problems

Worksheet 1. How do you prepare a 250.-ml of a 2.35 M HF dilution from a 15.0 M stock solution? 2. If 455-ml of 6.0 M HNO₃ is used to make a 2.5 L dilution, what is the molarity of the dilution? 3. If 65.5 ml of HCl stock solution is used to make 450.-ml of a 0.675 M HCl dilution, what is

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Getson's Blog

Molarity By Dilution

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Chemistry If8766 Fair

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Dilution Chemistry Pg

69 Answer molarity of

BaBr₂ solution:

$0.058375 \text{ mol} / 0.165 \text{ L}$

$= 0.35 \text{ M}$ Problem #9:

1.00 L of a solution is prepared by dissolving

125.6 Molarity By

Dilution 69 Answers -

Page 7/29 Molarity By

Dilution Chemistry Pg

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69 Answer

Worksheet

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Chemistry If8766

Answers

p. 31 Molarity By Dilution (mixing with water. a concentrated solvent) Acids are usually acquired from chemical supply houses in concentrated form. These acids are diluted to the desired concentration by adding water. Since moles of acid before

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dilution = moles of acid
after dilution, and
moles of acid : CV
then, $C_1 \times V_1 = C_2 \times V_2$.

Molarity By Dilution - Ms Beaucage

Answers Serial

Dilutions Practice

Worksheet Biol 307

Studocu . 1 if i have

340 ml of a 0.5 M NaBr

solution what will the

concentration be if i

add 560 ml more water

to it. Dilutions

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worksheet answer key.

Dilutions worksheet 1 if

i add 25 ml of water to

125 ml of a 0.15 M

NaOH solution what will

the molarity of the

diluted solution be.

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Services Program 1) If

45 mL of water are added to 250 mL of a 0.75 M K_2SO_4 solution, what will the molarity of the diluted solution be? 2) If water is added to 175 mL of a 0.45 M KOH solution until the volume is 250 mL, what

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filling and how much it

will cost to place the

filling in your mouth.

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Conc. Solution $M_1 \times V_1$,
= mol Water

Concentrated solution
is diluted with more
solvent. In this process,

$M_1 \times V_1 = \text{mol}$ USEFUL
EQUATIONS $M_1 \times V_1 =$

$M_2 \times V_2$ molarity =
mol solute / L. solution

1 L - 1000 ml. A solution
can be made less
concentrated in a
process called dilution.

This is accomplished
by adding more
solvent.

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**Dilution Name Chem
Worksheet 15-5**

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Solutions Worksheet 2
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Problems Answer Key
dissolved to make 0.10
L of solution. 2) 1.0
grams of potassium
fluoride is dissolved to
make 0.10 L of
solution. Molarity
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Answers

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And Dilution Problems

Answer Keyliters of

solution? 4.53 mol

$\text{LiNO}_3 = 1.59 \text{ M LiNO}_3$

3. 2.85 L soln Molarity

Worksheet 2 ANSWERS

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Problems Worksheet

$M=nV$ $n=$ # moles V

must be in liters

(change if necessary)

1. What is the molarity
of a 0.30 liter solution
containing 0.50 moles

Page 6/29

Solutions Worksheet

2 Molarity And

Dilution Problems ...

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Math Molarity Problems

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How to Calculate

Dilution of Solutions ...

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Solutions & Dilutions

Preparing solutions and
making dilutions

Simple dilutions Mixing
parts or volumes Serial

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dilutions Making fixed volumes of specific concentrations from liquid reagents:

$$(C1)(V1)=(C2)(V2)$$

Percent solutions (= parts per hundred)

Molar solutions

(unit=M=moles/L)

Lab Math Solutions, Dilutions, Concentrations and Molarity

3. What is the molarity of a solution of HNO_3 that contains 12.6

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grams HNO₃ in 1.0 L
of solution? ? mol HNO₃

$3 = 12.6 \text{ g HNO}_3 \times 1$

$\text{mol HNO}_3 \text{ } 63.0 \text{ g HNO}_3$

$3 = 0.200 \text{ mol HNO}_3 \text{ } 3 \text{ M}$

$= 0.200 \text{ mol HNO}_3 \text{ } 1.0$

$\text{L} = 0.200 \text{ M}$

4. How many grams of potassium nitrate are required to prepare 0.250 L of a 0.700 M solution? $0.700 \text{ M} =$ moles of solute 0.250 L moles of ...

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Work in groups on these problems. You should try to answer the questions without referring to your textbook. If you get stuck, try asking another group for help. Calculate molarity if 25.0 mL of 1.75 M HCl diluted to 65.0 mL.

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Calculate molarity by
dissolving 25.0g NaOH
in 325 mL of solution.

Worksheet

Answers

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