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Chapter 6: Chemical Kinetics 6-8 42.

Answers may vary. Sample answer: The rate of a chemical reaction is the change in ...

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Chapter 6 Review, pages 396-401

Chapter 4: Chemical Bonding: Unit 3:
Energy Changes and Rates of Reaction:
Chapter 5: Chemical Energy: Chapter 6:
Chemical Kinetics: Unit 4: Chemical
Systems and Equilibrium: Chapter 7:
Chemical Equilibrium: Chapter 8: Acid-
Base Equilibrium: Unit 5:
Electrochemistry: Chapter 9: Reduction-

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Oxidation Reactions: Chapter 10:
Electrochemical Cells

Nelson Ontario Senior Science Chemistry 12

6.4 Explore Applications of Chemical Kinetics: Biocatalysts and the Environment; 6.5 Rate Law; 6.6 Reaction Mechanisms; 6.7 Chemistry Journal:

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Global Warming and Reaction Rates;
Chapter 6 Investigations; Chapter 6
Summary; Chapter 6 Self-Quiz; Chapter
6 Review; Unit 3 Close. Unit 3 Task; Unit
3 Review; Unit 4: Chemical Systems and
Equilibrium

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Chemistry; Physical Chemistry; Chapter
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Courses. chem12_sm_07_5.pdf Size : 2054.636 Kb ... Size : 388.848 Kb Type : pdf Below are all of the resources for chapter 7 and 8. This is an important unit because there are a lot of questions on the exam and there are a lot of labs in this unit. ... 7.1 p. 420 in the ...

Pre University Courses

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2. Given: $r = 0.35 \text{ m}$; $N = 6.1 \times 10^6 \text{ C}$; $e = -1.602 \times 10^{-19} \text{ C}$; $k = 8.99 \times 10^9 \text{ N}\cdot\text{m}^2/\text{C}^2$
Required: ϵ ; V
Analysis:
Determine the charge on the object using $q = Ne$. Then calculate the magnitude of the electric field using $E = kq/r^2$ and the magnitude of the electric potential using $V = E \cdot d$.
Solution:
Determine the charge on the object:

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q=Ne =(6 ...

Section 7.6: The Millikan Oil Drop Experiment

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Chemistry 12 - Chapter 7 Quiz.

True/False. Indicate whether the sentence or statement is true or false. 1. Chemical equilibrium means that all chemical reactions have stopped. 2. Equilibrium can only occur in a closed

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system. 3. A catalyst shifts the position of equilibrium toward the products. ...

Chemistry 12 - Chapter 7 Quiz - Nelson

6.4 COLLISION THEORY AND RATE OF REACTION PRACTICE (Page 387)

Potential Energy, E_p Understanding Concepts 1. (a) Potential Energy

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Diagram for System R P P (i) (ii) (iii) R
Reaction Progress (b) The lower curve represents a catalyzed reaction; the upper curve represents the uncatalyzed reaction.

6.4 COLLISION THEORY AND RATE OF REACTION

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Chapter 6: Chemical Kinetics 6.2-1.

Section 6.2: Factors Affecting Reaction Rates. Section 6.2 Questions, page 365

1. (a) nature of reactant. (b) temperature.
(c) surface area. (d) nature of reactant.
(e) catalyst. (f) concentration.

**Section 6.2: Factors Affecting
Reaction Rates Section 6.2 ...**

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= 13.6 m/s [E] Required: p Analysis:
The momentum before the collision is equal to the momentum after the collision. Use the answer from (a) to determine the momentum after the collision. $p = (m_1 + m_2) v_f$ Solution:
 $p = (m_1 + m_2) v_f = (2200 \text{ kg} + 1300 \text{ kg})(13.6 \text{ m/s [E]})$
 $p = 4.8 \times 10^4 \text{ kg}\cdot\text{m/s}$
Statement: The momentum before and

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after the collision is ...

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