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(ANSWER 386.3g of LiNO<sub>3</sub>) 4) Using the following equation:  
$$\text{Fe}_2\text{O}_3 + 3 \text{H}_2 \rightarrow 2 \text{Fe} + 3 \text{H}_2\text{O}$$
. Calculate how many grams of iron can be made from 16.5 grams of Fe<sub>2</sub>O<sub>3</sub> by the following equation. Worksheet for Basic Stoichiometry. Part 1: Mole ↔ Mass Conversions. Convert the following number of moles of chemical into its corresponding

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mass in grams.

**Worksheet for Basic  
Stoichiometry**

Solution Stoichiometry  
Worksheet. Solve the  
following solutions

Stoichiometry

problems: 1. How

many grams of silver  
chromate will

precipitate when 150.

mL of 0.500 M silver

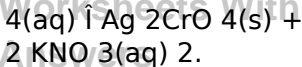
nitrate are added . to

100. mL of 0. 400 M

potassium chromate? 2

$\text{AgNO}_3(\text{aq}) + \text{K}_2\text{CrO}_4$

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**Solution**  
**Stoichiometry**  
**Worksheet**

Stoichiometry Practice  
Worksheet Solve the  
following stoichiometry  
grams-grams

problems: 1) Using the  
following equation:  $2 NaOH + H_2SO_4 \rightarrow 2 H_2O + Na_2SO_4$  How  
many grams of sodium  
sulfate will be formed if  
you start with 200.0



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grams of sodium hydroxide and you have an excess of sulfuric acid? 2) Using the following equation:

## **Stoichiometry Practice Worksheet**

Stoichiometry - Mass to Mass Problems Unit 6 Packet - Page 10 of 12 Stoichiometry Worksheet #1 Perform the following calculations. Be sure to use proper units!

Answer the following g

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- mol and/or mol  $\text{O}_3$  g  
conversion problems.
1. How many g in 7.00 mol of  $\text{N}_2$ ? \_\_\_\_\_ 2.  
How many g in 0.455 mol of  $\text{NaCl}$ ? \_\_\_\_\_ 3.  
How many mol in 23.0 g of  $\text{CaCO}_3$ ? \_\_\_\_\_

## **Unit 6 - Stoichiometry Packet**

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Stoichiometry: Mass to  
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**Mass To Mass**  
*Page 12/27*

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...

Stoichiometry  
WorkSheet #1: Worked  
Solutions Answer the  
following questions on  
your own paper. Show  
all work. Circle the final  
answer, giving units  
and the correct  
number of significant  
figures. 1. Based on  
the following equation,  
how many moles of  
each product are

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produced when 5.9  
moles of  $Zn(OH)_2$  are  
reacted with  $H_3PO_4$ ?  
(You need

## **Stoichiometry WorkSheet #1: Worked Solutions**

Stoichiometry  
Worksheet and Key  
1.65 mol  $KClO_3$  mol  
 $KClO_3$  mol  $O_2$  = mol  $O_2$   
2 3.50 mol  $KCl$  = mol  
 $KClO_3$  = 0.275 mol  $Fe$   
= mol  $Fe_2O_3$  = =

## **stoichiometry 1**

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**worksheet and key -  
Saddleback College**

CHM 130 Stoichiometry  
Worksheet KEY 1.

Fermentation is a complex chemical process of making wine by converting glucose into ethanol and carbon dioxide:  $C_6H_{12}O_6 (s) \rightarrow 2 C_2H_5OH (l) + 2 CO_2 (g)$  A.

Calculate the mass of ethanol produced if 500.0 grams of glucose reacts completely.

500.0 g  $C_6H_{12}O_6$  ..

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Answers

**CHM 130**  
**Stoichiometry**  
**Worksheet**

Stoichiometry  
Worksheet. In a  
combustion reaction,  
acetylene (C<sub>2</sub>H<sub>2</sub>)  
combines with oxygen  
to form carbon dioxide  
and water. Step 1:  
Write out the equation:  
\_\_\_ C<sub>2</sub>H<sub>2</sub> + \_\_\_ O<sub>2</sub> ( \_\_\_  
CO<sub>2</sub> + \_\_\_ H<sub>2</sub>O. 3.80  
moles of oxygen are



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used up in the reaction. How many moles of water are produced?

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1. Given the following equation:  $2 \text{C}_4\text{H}_{10} + 13 \text{O}_2 \rightarrow 8 \text{CO}_2 + 10 \text{H}_2\text{O}$ , show what the following molar ratios should be. a.  $\text{C}_4\text{H}_{10} / \text{O}_2$  b.  $\text{O}_2 / \text{CO}_2$  c.  $\text{O}_2 / \text{H}_2\text{O}$  d.  $\text{C}_4\text{H}_{10} / \text{CO}_2$  e.  $\text{C}_4\text{H}_{10} / \text{H}_2\text{O}$

2. Given the following equation:  $2 \text{KClO}_3 \rightarrow 2 \text{KCl} + 3 \text{O}_2$  a.

## Answers To Stoichiometry Worksheet

StoichiometryName

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Answers. 1. The human body needs at least  $1.03 \times 10^{-2}$  mol O<sub>2</sub> every minute. If all of this oxygen is used for the cellular respiration reaction that breaks down glucose, how many grams of glucose does the human bo

## **Stoichiometry - Difficult problems**

Step by Step:

Stoichiometry

Problems . Steps: 1)

Write the balanced

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chemical reaction. 2) Write a conversion equation. a) Find the mols of the compound with known mass. b) Use the mol ratio (in the balanced reaction) between the 2 compounds you are interested in. c) Find the grams of the compound you are looking for.

**Step by Step:  
Stoichiometry  
Problems Steps: Ex.**

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## **1) How ...**

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Stoichiometry is the tool for answering these questions.

Stoichiometry The study of quantitative relationships between the amounts of reactants used and amounts of products formed by a chemical reaction is called stoichiometry.

Stoichiometry is based on the law of

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conservation of mass.  
Recall from Chapter 3  
that the law states that

## **Chapter 11: Stoichiometry**

Examples and practice  
problems of solving  
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questions with gases  
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